



Europäisches Patentamt
European Patent Office
Office européen des brevets



⑪ Publication number:

0 596 515 A1

⑫

EUROPEAN PATENT APPLICATION

㉑ Application number: 93117966.7

㉓ Int. Cl.5: G03F 7/42

㉒ Date of filing: 05.11.93

㉔ Priority: 06.11.92 US 972511

㉕ Applicant: J.T. BAKER INC.

㉖ Date of publication of application:
11.05.94 Bulletin 94/19

222 Red School Lane
Phillipsburg New Jersey 08865(US)

㉗ Designated Contracting States:
BE CH DE ES FR GB IE IT LI LU NL SE

㉘ Inventor: Schwartzkopf, George
Bickel Road
Franklin Township, New Jersey 07882(US)

㉙ Representative: VOSSIUS & PARTNER
Siebertstrasse 4
D-81675 München (DE)

㉚ Alkaline photoresist stripping composition producing reduced metal corrosion.

㉛ Addition of non-nitrogen containing weak acids to amine-containing alkaline strippers for photoresists produce stripper compositions able to strip highly cross-linked or hardened photoresist films without producing any substantial metal corrosion. Weak acids useful in the stripping compositions include those having a pK in aqueous solution of 2.0 or higher and an equivalent weight of less than about 140 and are employed in an amount to neutralize from about 19% to about 75% of the amine present in the stripper composition.

EP 0 596 515 A1

This invention relates to alkaline-containing photoresist stripping compositions that avoid or substantially eliminate metal corrosion on microcircuit substrates containing metals and coated with cross-linked or hardened photoresist resin. This invention also relates to a process for stripping photoresist from cross-linked or hardened substrates employing alkaline-containing photoresist stripping compositions without incurring any substantial metal corrosion in microcircuit substrates containing metal and which provide greatly improved strip rates, with consequent decrease in strip time, compared to stripping compositions not containing the alkaline component.

5 An integral part of microelectronic fabrication is the use of photoresists to transfer an image from a mask or reticle to the desired circuit layer. After the desired image transfer has been achieved, the photoresist is removed by stripping before proceeding to some subsequent process step. Since about 1980, amides and mixtures of amides with various co-solvents have routinely been used for this stripping step, see for example U.S. 4,395,479; U.S. 4,428,871 and U.S. 4,401,748, issued to Ward et al.

10 Depending upon the type of pre-stripping processing that is performed while the photoresist is in place, the photoresist polymer may be cross-linked or hardened to the extent that amide-based solvents will no longer effectively strip the photoresist. Since about 1985, the use of amide mixtures containing additional 15 alkaline components, such as organic amines (Johnson U.S. 4,592,787; Merrem U.S. 4,765,844; Sizensky U.S. 4,617,251; Turner WO 87/05314; Thomas et al. U.S. 4,791,043) or quaternary ammonium hydroxides (Steppan et al. U.S. 4,776,892; Hag U.S. 4,744,834; Martin WO 88/05813) were introduced to facilitate the removal of such hardened photoresists.

20 The use of these alkaline strippers on microcircuit substrates containing metal films, particularly aluminum or various combinations or alloys of active metals such as aluminum or titanium with more electropositive metals such as copper or tungsten, has proven problematic. Various types of metal corrosion, such as corrosion whiskers, pitting, notching of metal lines, have been observed due, at least in part, to reaction of the metals with alkaline strippers. Further it has been shown, by Lee et al., Proc. 25 Interface '89, pp. 137-148, that very little corrosive action takes place until the water rinsing step that is required to remove the organic stripper from the wafer. The corrosion is evidently a result of contacting the metals with the strongly alkaline aqueous solution that is present during rinsing. Aluminum metal is known to corrode rapidly under such conditions - Ambat et al., Corrosion Science, Vol. 33 (5), p. 684, 1992.

30 In the past it has been proposed to circumvent this problem by employing intermediate rinses with non-alkaline organic solvents such as isopropyl alcohol. It would be desirable to avoid the expense and possible safety, chemical hygiene, and environmental consequences of such an intermediate rinse by eliminating this additional rinse by providing an effective stripper that does not require it.

In U.S. Patent 5,102,777, issued on April 27, 1992 to Lin et al., it is proposed to combat such corrosion by employing a positive photoresist stripper composition which includes a solvent system having solubility 35 parameters which fall within a range of from about 8.5 to about 15 in an amount which falls within a range of about 65% to about 98%. In addition, an amine is present in an amount of from about 2% to about 25% and also a fatty acid having 8 to 20, preferably 10 to 16, carbon atoms in an amount of from about 0.1% to about 10% such that the amount of amine and fatty acid is selected to provide a pH of from about 6 to 9.5, preferably 7 to 8.5. However, while the positive photoresist stripping compositions disclosed in said patent 40 were found to provide some beneficial effect vis-a-vis metal corrosion, these disclosed stripping compositions were not able to be used to satisfactorily strip hard-to-strip cross-linked or hardened photoresist resins. Moreover, in the one stripping formulation of the patent which is capable of stripping cross-linked photoresist resins, one is not able to prevent or avoid metal corrosion. These deficiencies and drawbacks of the stripping compositions of the Lin et al. patent are demonstrated in the Comparative Example set forth 45 hereinafter in this specification.

It is therefore an object of this invention to provide alkaline-containing photoresist strippers which do not require intermediate rinses to avoid metal corrosion and yet the photoresist strippers are still highly effective in stripping cross-linked or hardened photoresists. A further object of this invention is to provide such improved non metal corroding, alkaline-containing photoresist compositions without any undue adverse 50 effect on strip rate of the photoresist for cross-linked or hardened photoresists.

These objects are achieved by an alkaline-containing photoresist stripping composition able to strip hardened or highly cross-linked photoresist resin film from a substrate containing metal without producing any substantial metal corrosion comprising a stripping solvent, a nucleophilic amine and a non-nitrogen containing weak acid in an amount sufficient to partially neutralize the nucleophilic amine.

55 The addition of certain non-nitrogen containing weak acids to alkaline strippers results in stripper compositions that yield considerably less alkaline solution during rinsing with pure water but nevertheless retain their ability to strip highly cross-linked or hardened photoresist films without producing any substantial undesirable metal corrosion. This is especially surprising since more than half of the alkaline

component can be neutralized with the weak acid without significantly adversely altering the stripping times or rates. It was also found that these lower pH aqueous solutions cause considerably less corrosion on metal than the water rinse solutions resulting from alkaline-containing strippers that are not modified via the weak acid addition of this invention.

- 5 Furthermore, the photoresist strippers of this invention are able to satisfactorily strip hardened or cross-linked photoresist resins while producing substantially no undesirable metal corrosion.

The non-nitrogen containing weak acids that may be employed in this invention include organics such as carboxylic acids or phenols as well as salts of inorganic acids such as carbonic or hydrofluoric acid.

- 10 By weak acids is meant acids having a strength expressed as a "pK" for the dissociation constant in aqueous solution of at least 2.0 or higher, preferably 2.5 or higher. Particularly useful are weak acids of pK > 2.0 and preferably having an equivalent weight of less than about 140. As examples of such non-nitrogen containing weak acids useful in this invention there may be mentioned, for example, carboxylic acids such as acetic acid, phthalic acid and phenoxyacetic acid, organic acids such as e.g. 2-mercaptopbenzoic acid and 2-mercaptoethanol, phenols generally having pK in the range of from 9 to 10, such as e.g. phenol, 15 1,3,5-trihydroxybenzene, pyrogallol, resorcinol and 4-tert-butylcatechol, and inorganic acids such as e.g. carbonic acid and hydrofluoric acid.

- 20 The amount of weak acid employed in the stripping compositions of this invention is from about 0.05% to about 25% by weight of said composition so that the amount is sufficient to neutralize about 19% to about 75% by weight of the amine present in the stripper composition thereby resulting in an aqueous rinse pH for said stripper compositions of from about pH 9.6 to about 10.9.

- 25 Alkaline stripper components that may be used in this invention also cover a wide range of structural types. Their dissociation constants, once again expressed as pK values, range from about 9 to 11 for the beta-oxygen or -nitrogen substituted amines to 8.3 for the secondary amine, morpholine and hydroxylamines and hydroxylamine derivatives of somewhat lower pK values. Among the alkaline components that may be used there may be mentioned, nucleophilic amines, preferably for example, 1-amino-2-propanol, 2-(2-aminoethoxy)ethanol, 2-aminoethanol, 2-(2-aminoethylamino)ethanol and 2-(2-aminoethylamino)ethylamine.

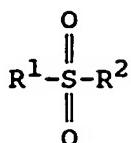
- 30 More important than the actual pK value of an amine is its nucleophilicity which should be high. The amount of amine component employed in the stripping compositions of this invention is from about 1% to about 50% by weight of said composition.

It is believed that the interaction of these alkaline stripper components with the range of weak acids used in this invention is essentially reversible:



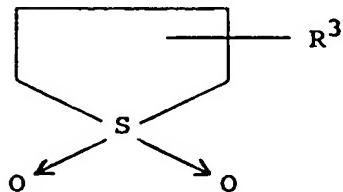
- Because of the reversibility of this reaction, substantial concentrations of the alkaline component would remain available during the stripping process even though much of the alkaline component has been neutralized from a stoichiometric point of view. This would account for the surprisingly rapid stripping rates that are observed even in the presence of these acids.

- 40 The photoresist stripping compositions of this invention which contain the aforementioned alkaline component and weak acid component will also comprise an organic solvent system. The organic solvent system is one having a solubility parameter of from about 8 to about 15, obtained by taking the square root of the sum of the squares of the three Hansen solubility parameters (dispersive, polar and hydrogen bonding). The solvent system may comprise any of a number of individual solvents or a mixture of several different solvents. As example of such solvents there may be mentioned, various pyrrolidinone compounds such as e.g. 2-pyrrolidinone, 1-methyl-2-pyrrolidinone, 1-ethyl-2-pyrrolidinone, 1-propyl-2-pyrrolidinone, 1-50 hydroxyethyl-2-pyrrolidinone and 1-hydroxypropyl-2-pyrrolidinone, diethylene glycol monoalkyl ethers such as those of the formula HOCH₂CH₂-O-CH₂CH₂-O-R where R is an alkyl radical of from 1 to 4 carbon atoms, compounds containing sulfur oxides such as dialkyl sulfones of the formula



wherein R¹ and R² are alkyl of 1 to 4 carbon atoms, dimethyl sulfoxide (DMSO), tetrahydrothiophene-1,1-dioxide compounds of the formula

10



20

wherein R³ is hydrogen, methyl or ethyl, such as sulfolane, methyl sulfolane and ethyl sulfolane, as well as polyethylene glycols, dimethylacetamide or dimethylformamide. The solvent system portion of the stripper compositions of this invention will generally comprise from about 50% to about 98% by weight (which encompasses the range of 49.95% to 98% by weight) of the composition, preferably about 85% to about 25 98% by weight.

25

The stripping compositions of this invention are effective in stripping a wide and varied range of photoresists especially positive photoresists. Most photoresists consist of an ortho naphthoquinone diazide sulfonic acid ester or amide sensitizer or photoactive component, with novolak, resole, polyacrylamide or acrylic copolymer type binders or resins. Such photoresists are well known in the art. Such resists and sensitizers are described for example, in U.S. Patent Nos. 3,046,118; 3,046,121; 3,106,465; 3,201,239; 30 3,538,137; 3,666,473; 3,934,057; 3,984,582 and 4,007,047. As examples of photoresist compositions for which the stripping composition of this invention may be used there may be mentioned KTI Chemicals photoresist KTI 820, KTI 825, KTI 875 and KTI 895i; J. T. Baker Inc. photoresist 1-PR-21 and E38; Olin Hunt WX-309, HiPR 6500 series photoresists and OiR 3712 photoresist; Hoechst Celanese photoresists 5214-E, 35 AZ-5200 and AZ-6212; Shipley Company photoresists MF-314, XP-8843, Megaposit SPR 500 and Megaposit SNR 248; Fuji Hunt Electronics photoresists FH 6450 and FX-EX1; Tokyo Ohka Kogyo Co. Ltd. photoresist THMR-iP800; MacDermid photoresist PR-1024 MB; Dynachem Division of Morton International photoresists Nova 2000 series, e.g. Nova 2020 and Nova 2070; Japan Synthetic Rubber photoresist PFR 1X500EL; Sumitomo Chemical photoresist Sumiresist PFI-15; and Toray Industries photoresist PR-a1200, 40 The examples illustrate the invention.

EXAMPLE 1

Silicon wafers were primed with hexamethyldisilazane and spun with about 1000 nm of J. T. Baker Inc. 45 1-PR-21 photoresist. The wafers were then soft-baked for 30 min. at 95 °C, then hard-baked for 60 min. at 200 °C giving a highly cross-linked resist film. The wafers were then stripped at 95 °C in the solutions listed below and the time at which the photoresist layer was fully stripped was recorded. These are shown below (plus or minus one standard deviation of the several replicate wafers that were used). The alkaline stripper used in this example comprised 90% of N-methylpyrrolidinone (NMP) and 10% of 1-amino-2-propanol. The 50 acid used was acetic acid. Water rinse pH was determined at 19:1 dilution of these mixtures with deionized water and is also listed. The aluminum (foil coupons = 1.5 x 1.5 x .004 inches) corrosion weight loss shown was determined at 9:1 dilution with deionized water at room temperature for 24 hours.

55

EP 0 596 515 A1

	Formulation	Gm. acid/kg.formulation	Strip Time (sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
5	NMP only	0	390 ± 50			
	Alkaline stripper	0	107 ± 5	0	11.6	13
	Alkaline stripper	10	96 ± 2	12.5	10.8	5
	Alkaline stripper	20	104 ± 9	25	10.4	2
	Alkaline stripper	30	98 ± 3	37.5	10.2	1
	Alkaline stripper	40	98 ± 7	50	9.9	0.9
	Alkaline stripper	50	103 ± 4	62.5	9.7	0.4

EXAMPLE 2

15 This example was conducted according to the procedure of Example 1 except that the alkaline stripper comprised 90% of dimethylformamide (DMF) and 10% of 2-(2-aminoethoxy)ethanol, the acid was phenol, and the stripping temperature was 85 °C.

	Formulation	Gm.acid/kg.formulation	Strip Time (sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
20	DMF only	0	426 ± 44			
	Alkaline stripper	0	47 ± 2	0	11.2	9.4
	Alkaline stripper	10	48 ± 2	11	10.8	4.8
	Alkaline stripper	20	52 ± 1	22	10.6	3.4
	Alkaline stripper	30	54 ± 1	33	10.5	3.8
	Alkaline stripper	40	55 ± 1	44	10.4	2.1
	Alkaline stripper	50	55 ± 2	55	10.3	2.4

30

EXAMPLE 3

35 This example was conducted according to the procedure of Example 1 except that the alkaline stripper comprised 90% of dimethylacetamide (DMAc) and 10% of 2-aminoethanol, the acid was 1,3,5-trihydroxybenzene, and the stripping temperature was 85 °C.

	Formulation	Gm.acid/kg.formulation	Strip Time (sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
40	DMAc only	0	343 ± 34			
	Alkaline stripper	0	51 ± 7	0	11.7	16
	Alkaline stripper	11	56 ± 2	12.5	11.0	11
	Alkaline stripper	22	55 ± 5	25	10.7	9
	Alkaline stripper	33	53 ± 3	37.5	10.6	5
	Alkaline stripper	44	53 ± 3	50	10.4	4
	Alkaline stripper	55	61 ± 2	62.5	10.3	2

50 **EXAMPLE 4**

55 This example was conducted according to the procedure of Example 1 (alkaline stripper = 90% NMP and 10% 1-amino-2-propanol) except that the acid component was varied through a series of weak acids of various structural types. Stripping temperature was 85 °C.

	Formulation	Gm.acid/kg.formulation	Strip Time (sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
5	NMP only	0	648			
	Alkaline stripper(AS)	0	219 ± 31	0	11.6	13
	AS + pyrogallol	50	231 ± 15	69	10.3	5
	AS + resorcinol	50	251 ± 19	68	10.4	4
	AS + 4-tert-butylcatechol	70	245 ± 29	63	10.5	0
	AS + phthalic acid	70	238 ± 36	63	9.7	0.7
	AS + phenoxyacetic acid	100	257 ± 22	49	9.9	1

EXAMPLE 5

15

This example was conducted according to the procedure of Example 1 except that two different alkaline stripper systems were used: AS 1 = 90% dimethylacetamide (DMAc) + 10% 2-(2-aminoethylamino)ethanol with or without added acetic acid; AS 2 = 90% dimethylacetamide (DMAc) + 6% 2-(2-aminoethylamino)-ethylamine with or without added phenol. Stripping temperature was 85 °C.

20

	Formulation	Gm.acid/kg.formulation	Strip Time (sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
25	DMAc only	0	471 ± 38			
	AS 1	0	97 ± 2	0	11.4	8
	AS 1 + acetic acid	4	149 ± 10	69	9.6	0.4
	AS 2	0	242 ± 48	0	11.4	8
	AS 2 + phenol	6	226 ± 26	53	10.3	3

EXAMPLE 6

35

This example was conducted according to the procedure of Example 1 except that the alkaline stripper comprised 90% of 2-pyrrolidinone and 10% 1-amino-2-propanol, the acid was acetic acid, and the stripping temperature was 100 °C.

45

	Formulation	Gm.acid/kg.formulation	Strip Time (sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
40	2-pyrrolidinone only	0	491 ± 14			
	Alkaline stripper	0	201 ± 9	0	11.5	14
	Alkaline stripper	50	221 ± 24	62.5	9.6	0

EXAMPLE 7

50

This example was conducted according to the procedure of Example 1 except that three different alkaline stripper systems were used and the pH controlling materials were weak inorganic acids: AS 1 = 90% NMP + 10% 1-amino-2-propanol with or without added carbonic acid (generated via ammonium carbonate addition); AS 2 = 90% NMP + 10% 2-(2-aminoethoxy)ethanol with or without added 47% aq. hydrofluoric acid; AS 3 = 90% NMP + 10% 2-aminoethanol with or without added 47% aq. hydrofluoric acid. Stripping temperature was 85 °C.

55

EP 0 596 515 A1

	Formulation	Gm.acid/kg.formulation	Strip Time (sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
5	NMP only	0	337 ± 36	0	11.4	10
	AS 1	0	135 ± 12	0	10.0	2
	AS 1 + carbonic acid	26	126 ± 16	62	11.3	9
10	AS 2	0	132 ± 12	0	9.8	2
	AS 2 + HF	9	110 ± 4	49	11.5	17
	AS 3	0	67 ± 3	0	10.4	3
	AS 3 + HF	16	70 ± 9	50		

EXAMPLE 8

In a manner similar to Example 4, the acid component was varied through a series of sulfur-containing acid types. In this example the alkaline stripper was 90% of N-methylpyrrolidinone (NMP) and 10% of 2-aminoethanol and the photoresist used was commercially available, KTI 895i photoresist. Processing was similar to the previously used Baker 1-PR-21 including cross-linking at 200°C for 60 minutes. A stripping temperature of 95°C was utilized:

	Formulation	Gm.acid/kg.formulation	Strip Time (sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
25	NMP only	0	359 ± 30	0	11.6	14
	Alkaline stripper(AS)	0	48 ± 6	0	10.2	0.5
	AS + 2-mercaptopbenzoic acid	60	86 ± 3	48	10.4	2
30	AS + 2-mercaptoethanol	60	49 ± 3	47		

EXAMPLE 9

35 This example was conducted in a manner similar to Example 1. The alkaline stripper was 90% N-methylpyrrolidinone (NMP) and 10% 1-amino-2-propanol plus 50 gm. acetic acid/kg. formulation. The effect of adding 10% by weight of various common stripper co-solvents was measured. Stripping temperature was 85°C.

	Formulation	Identity of 10% co-solvent	Strip Time(sec)	% Amine neutralized	Water pH	Corrosion weight loss (%)
40	NMP only	none	475 ± 38			
45	Alkaline stripper (AS)	none	219 ± 21	62.5	9.7	0.7
	AS + sulfolane		290 ± 4	62.5	9.7	1
	AS + N-(2-hydroxyethyl) pyrrolidinone		254 ± 11	62.5	9.7	0.9
	AS + N-cyclohexylpyrrolidinone		297 ± 17	62.5	9.7	0.4
50	AS + triethyleneglycol methyl ether		266 ± 9	62.5	9.6	0.9
	AS + diethyleneglycol ethyl ether		283 ± 2	62.5	9.6	1

This example shows that co-solvents may be added to these partially neutralized compositions with retention of the desired pH and corrosion properties and only a slight loss of stripping speed.

EXAMPLE 10

This example was conducted in a manner similar to Example 1 except that the alkaline stripper employed 90% dimethyl sulfoxide (DMSO) solvent and 10% of 1-amino-2-propanol. The acid employed was

acetic acid. Stripping temperature was 80°C.

Formulation	Gm. acid/kg.formulation	Strip Time(sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
DMSO only	0	356 ± 27			
Alkaline stripper	0	45 ± 5	0	11.6	11
Alkaline stripper	10	53 ± 9	12.5	10.7	4
Alkaline stripper	20	60 ± 9	25	10.4	2
Alkaline stripper	30	84 ± 8	37.5	10.1	0.7
Alkaline stripper	40	70 ± 10	50	9.9	0.4
Alkaline stripper	50	93 ± 29	62.5	9.7	0.3

COMPARATIVE EXAMPLE

This is a comparative example of the six stripper compositions exemplified in column 5 of U.S. 5,102,777. The table in U.S. 5,102,777 does not specify the processing conditions or the photoresist used. In this comparative example the cross-linked photoresist resin KTI 895i was used and baked at 200°C for 60 minutes and stripped at a stripping temperature of 85-90°C, which approximates the degree of photoresist removal difficulty that is commonly encountered in current integrated circuit fabrication and is typical of the conditions used in the examples of this invention. An example of a stripping composition with only N-methylpyrrolidinone is included as a basis for comparison.

Solvent	Acid	Amine	Strip Time(sec)	% Amine neutralized	Water rinse pH	Corrosion weight loss (%)
100% NMP	none	none	> 900b			
95% NMP	.5% capric	4.5% TEA	> 900a	10	9.1	0.7
95% DE	.5% capric	4.5% TEA	>1800b	10	8.8	
50% NMP/46.5% DE	.5% capric	3% TEA	>1800b	14	8.7	
50% NMP/46.5% BLO	.5% lauric	3% DGA	>1800b	*	water insoluble	
50% NMP/46.5% DE	.5% lauric	3% DEA	>1800b	9	10.0	
70% NMP/25% DMAc	.5% lauric	4.5% MPA	92a	4	11.5	16
69.8% NMP/25% DMAc	.7% acetic	4.5% MPA	127a	19	11.1	7
94.8% NMP	.7% acetic	4.5% MIPA	188a	19	10.6	3
87.4% NMP	2.9%acetic	9.7% MIPA	185b	38	10.2	1

NMP = N-methylpyrrolidinone
DE = diethyleneglycol monobutyl ether
BLO = gamma-butyrolactone
DMAc = dimethylacetamide
TEA = triethanolamine
DGA = diglycolamine ([2-aminoethoxy]-ethanol)
DEA = diethanolamine
MPA = monopropanolamine (3-aminopropanol)
MIPA = monoisopropanolamine (1-amino-2-propanol)
* BLO reacts with DGA at room temperature forming an amide, thus no amine is present to neutralize

a = strip temperature of 85°C

b = strip temperature of 90°C.

Evidently rather mild processing conditions are in use for the examples in U.S. 5,102,777 (center section of the foregoing data) since the first five of said stripping compositions do not strip cross-linked

photoresists and are essentially equivalent to unmodified NMP in stripping ability. The only stripping composition example of U.S. 5,102,777 capable of removing this KTI 895i resist (the sixth entry in the center section of the foregoing data) also caused considerable aluminum corrosion (16% weight loss). This is due to insufficient neutralization (4%) of the amine present resulting in a pH of 11.5.

- 5 The last three stripper examples of the foregoing table are within the scope of the present invention, which demonstrates the ability of the stripping compositions of this invention to strip cross-linked photoresists with reduced levels of aluminum corrosion.

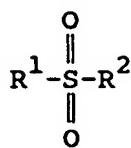
In each of the foregoing examples of stripping compositions according to this invention, addition of the alkaline material to the solvent decreases the strip time (increases the strip rate) markedly demonstrating that alkaline strippers are effective for highly cross-linked photoresists. However, neutralization of a substantial portion of these alkaline materials, over the broad range of solvent, base, and acid types utilized in Examples 1-10, unexpectedly gives new formulations that retain this property of stripping cross-linked resists. At the same time these formulations yield rinse water of considerably lower pH. The lower rinse pH resulting from these modified photoresist stripper formulations gives dramatically less aluminum metal corrosion. Other metals and metal combinations that are subject to alkaline corrosion would be expected to show similar improved behavior.

With the foregoing description of the invention, those skilled in the art will appreciate that modifications may be made to the invention without departing from the spirit thereof. Therefore, it is not intended that the scope of the invention be limited to the specific embodiments illustrated and described.

20

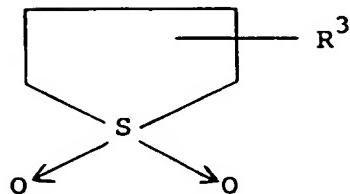
Claims

1. An alkaline-containing photoresist stripping composition able to strip hardened or highly cross-linked photoresist resin film from a substrate containing metal without producing any substantial metal corrosion comprising a stripping solvent, a nucleophilic amine and a non-nitrogen containing weak acid in an amount sufficient to partially neutralize the nucleophilic amine.
2. An alkaline-containing photoresist stripping composition according to Claim 1 wherein:
the stripping solvent is a stripping solvent system having a solubility parameter of from about 8 to about 15 and is present in an amount of from about 50% to about 98% by weight of the stripping composition;
the nucleophilic amine is present in an amount of from about 1 to about 50 % by weight of the stripping composition; and
the weak acid has a pK value in aqueous solution of 2.0 or greater and an equivalent weight of less than 140 and is present in the stripping composition in an amount of from about 0.05 % to about 25 % by weight of said stripping composition, so that the amount is sufficient to neutralize from about 19 % to about 75 % by weight of the nucleophilic amine such that the stripping composition has an aqueous pH of within the range of from about pH 9.6 to about 10.9.
3. An alkaline-containing photoresist stripping composition according to any one of Claims 1 or 2 wherein the weak acid has a pK of 2.5 or higher and is selected from the group consisting of acetic acid, phthalic acid, 2-mercaptopbenzoic acid, 2-mercaptoethanol, 1,3,5-trihydroxybenzene, pyrogallol, resorcinol, 4-tert-butylcatechol, carbonic acid and hydrofluoric acid, and/or the amine is selected from the group consisting of 1-amino-2-propanol, 2-(2-aminoethoxy)ethanol, 2-aminoethanol, 2-(2-aminoethylamino)ethanol and 2-(2-aminoethylamino)ethylamine.
4. An alkaline-containing photoresist stripping composition according to any one of Claims 1 to 3 wherein the stripping solvent comprises one or more solvents selected from the group consisting of 2-pyrrolidinone, 1-methyl-2-pyrrolidinone, 1-ethyl-2-pyrrolidinone, 1-propyl-2-pyrrolidinone, 1-hydroxyethyl-2-pyrrolidinone, 1-hydroxypropyl-2-pyrrolidinone, diethylene glycol monoalkyl ethers of the formula $\text{HOCH}_2\text{CH}_2\text{-O-CH}_2\text{CH}_2\text{-O-R}$ where R is an alkyl radical of from 1 to 4 carbon atoms, dialkyl sulfones of the formula



where R¹ and R² are alkyl of 1 to 4 carbon atoms, dimethyl sulfoxide, tetrahydrothiophene-1,1-dioxide compounds of the formula

10



20

wherein R³ is hydrogen, methyl or ethyl, a polyethylene glycol, dimethylacetamide and dimethylformamide.

25

5. An alkaline-containing photoresist stripping composition according to any one of Claims 1 to 4 wherein said composition comprises on a weight basis from about 50% to about 98% by weight of N-methylpyrrolidinone, from about 1 to about 50% of an amine selected from 1-amino-2-propanol, 2-(2-aminoethoxy)ethanol and 2-aminoethanol and from about 0.05% to about 25% of acetic acid present in an amount whereby the amine is neutralized such that the stripping composition has an aqueous pH within the range of from about pH 9.6 to about 10.9.
- 30 6. A method of stripping highly cross-linked or hardened photoresist resin from a substrate with a photoresist stripping composition using the alkaline-containing photoresist stripping composition according to any one of Claims 1 to 5 as the photoresist stripping composition.

35

40

45

50

55



European Patent
Office

EUROPEAN SEARCH REPORT

Application Number

EP 93 11 7966

DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int. CL.5)
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
X	US-A-5 102 777 (WEI-YUAN LIN ET AL.) 7 April 1992 * claims 1,2,12,19,21,27 * ---	1-6	G03F7/42
X	EP-A-0 466 071 (KONICA CORPORATION) 15 January 1992 * claim 1; examples 1-3 * ---	1-6	
X	IBM TECHNICAL DISCLOSURE BULLETIN vol. 14, no. 10, 1 March 1972, USA page 3132 'Removing Silicone Gel from Semiconductor Multichip Modules' ---	1-6	
Y	DE-A-4 103 586 (FUJI PHOTO FILM CO. LTD.08.08.91) * page 3, line 13 * * page 3, line 34 - line 48; claim 1 * ---	1-6	
Y	DE-A-2 823 414 (EASTMAN KODAK CO.) 14 December 1978 * claims 1,9; example 1 * ---	1-6	
A	IBM TECHNICAL DISCLOSURE BULLETIN vol. 32, no. 3A, 1989, USA , XP49420 'Acetic acid Spray Rinse for removing, developed Resist from a Wafer Surface' -----	1-6	G03F TECHNICAL FIELDS SEARCHED (Int. CL.5)
<p>The present search report has been drawn up for all claims</p>			
Place of search MUNICH	Date of completion of the search 02 MARCH 1994	Examiner Thiele N.K.	
CATEGORY OF CITED DOCUMENTS		T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons ----- & : member of the same patent family, corresponding document	
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document			